



PROCESS FOR GENERATION OF PRECIPITATED CALCIUM CARBONATE FROM CALCIUM CARBONATE RICH INDUSTRIAL BY-PRODUCT

5 Field of the Invention

The present invention relates to a process for the production of precipitated calcium carbonate from a calcium carbonate-rich by-product generated in an industrial processes, specifically, a by-product generated by a nitrophosphate plants in the fertilizer industry.

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Background of the Invention

TC 1700

Precipitated calcium carbonate is produced by calcination of calcium carbonate rich by-product into quick lime in a kiln, followed by slaking or hydration to obtain hydrated lime and subsequently subjecting it to carbonation.

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Precipitated calcium carbonate finds varied commercial uses like, in, including the manufacture of paper, rubber, plastics, glass, textiles, putties, chalks, sealant, adhesives, paints, inks, varnishes, food, cosmetics, dentrifices, chemicals and pharmaceuticals.

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Commercial applications of precipitated calcium carbonate require well-defined powder characteristics, particularly, fine particles with a narrow size distribution, uniform shape and crystallinity. Marentette J.M. *et al.* ("Crystallisation of Calcium Carbonate in the presence of PEO-block-PMAA copolymers", *Adv. Mater.*, 9, 647, 1997) have shown that these characteristics play a crucial role in product properties and that their control is important for the preparation of industrially useful products. The precipitated calcium carbonate must also needs to be substantially free of impurities for being to be useful for various commercial applications. Several physical and chemical processes have been reported for the treatment of solid waste containing calcium carbonate.

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Physical processes requires drying and grinding to such a fineness that allows impurities can to be removed by screening, classification, magnetic separation, hydrocyclone and floatation separation. The disadvantages of these physical methods are include the requirement of special

~~machines/devices/equipment~~ and its—the required maintenance thereof. Other disadvantages of existing processes include The unpredictable process efficiencies, variation—variable in results of—in the quantities of impurities removed and the expenses associated with employing such physical techniques—are other
5 disadvantages.

Chemical processes for the purification of calcium carbonate-rich waste involves leaching or bleaching of impurities using special reagents. Disadvantages of these methods are that itinclude the require requirement of a number of unit operations to perform the treatment. Moreover, a single chemical process may not
10 be suitable to—for remove removal of all the impurities. Removal and separation of excess chemical reagents after the treatment is another disadvantage.

Another previously known and widely used method for the treatment of calcium carbonate-rich waste is reburning within the kiln and removing it—the reburned waste after cooling so—as to obtain calcium oxide. The calcium oxide
15 which is then recycled in a causticizing process in producing paper pulp or it is subjected to hydration followed by carbonation to produce precipitated calcium carbonate.

Reference may be made to US patent No. 4,018,877 (1977) to Richard Derek Anthony Woode according to which In the process of Richard Woode (US
20 Patent 4,018,877), an aqueous suspension of calcium hydroxide at 25°C is—was agitated vigorously and reacted with a mixture of air and carbon dioxide. After 15 minutes (following the ‘primary nucleation stage’) a complex-forming agent—for ealeium ions, such as a hydroxy carboxylic acid, especially—particularly a hydroxy poly-carboxylic acids (for example, e.g., Citrie—citric acid and malic acid) which
25 complexes calcium ions. The complexing agent was added in athe concentration range of 0.001 to 5 wt.%, especially—preferably in the range 0.03 to 0.2 wt.% by weight based on the weight of the calcium carbonate produced—is—added. The carbonation was stopped after about—a further 50 minutes when the reaction mixture had just become acid to a phenolphthalein indicator. The mixture is—was then
30 heated to 85°C over a period of 20 minutes and is—was allowed to age for 30 minutes. Carbonation is—was restarted at the—a much lower rate, maintaining the temperature to—at 85°C, and aAfter 20 to 40 minutes the pH of the batch had fallen

below 8.0. At this stage, 0.8% of stearic acid in ammoniacal solution is-was added and the mixture is-was stirred at 85°C for about 3 hours. The suspension is-was filtered, and-The filter cake is-was extruded through 5/16 inch diameter holes to yield "granules" which are-were dried in an oven overnight at 130°C on a gauze-tray to produce calcium carbonate having 0.72 relative granule hardness and 0.07 micron ultimate particle size with a soft texture. The drawbacks of this process are that the total batch/production time is more than 5 hours during which time the temperature is maintained at 85°C for a period of 4 hours, and-In addition, the process requires drying of the product overnight at 130°C-for overnight period.

5 This makes this process is thus highly energy consuming and is therefore unattractive.

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In US Patent No. 4,133,894 (1979) to Hiroji Shibasaki *et al.* (US Patent 4,133,894), it is disclosed disclose that the precipitates of uniform particle size can be continuously produced by repeating the step of carbonation reaction. In the first 15 step of the process, a suspension of calcium hydroxide having a solids concentration 0.1 to 10 weight % and a temperature of 15 to 30°C is sprayed in the form of droplets of about 0.2 to 1.0 mm in diameter against a gas containing 10 to 40 volume % carbon dioxide in countercurrent contact therewith, the-The gas being-is passed at a specified superficial velocity of about 0.02 to 0.5 m/sec., whereby By this process, 5 to 15% of the calcium hydroxide is converted to calcium carbonate. In the second step of this process, the suspension resulting from the first step is sprayed in the form of droplets of about 1.0 to 1.5 mm diameter against a gas containing 15 to 35 volume % of carbon dioxide and passed upward through the column at a superficial velocity of about 1.5 to 2.5 m/sec whereby growth of the crystals is accomplished. In the third step of this process, the suspension resulting from second step is sprayed at a temperature of up to 30°C and in the form of droplets of about 1.5 to 2.0 mm in diameter into a column in countercurrent contact at a superficial velocity of about 1.5 to 3.0 m/sec. whereby the carbonation is completed. Thus, superfine calcium carbonate having an 25 average particle size of less than about 0.1 to 3.0 microns is produced. The main drawback of this invention process is that it needs-requires control of number of parameters such as solids concentration, droplet size, temperature of suspension,

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gas velocity of carbon dioxide containing gas etc. at a time for three columns. Another drawback is use the requirement of multi-step- carbonation which is more expensive in terms of operating cost for columns and pumps than a require much capital investment for columns and pumps and the operating cost, compared to single stage carbonation.

US patent 5,342,600 (1994) issued to Bleakley, Ian S. *et al.* (US Patent 5,342,600) describes a method of preparing precipitated calcium carbonate which comprises; (1) slaking quick lime in an aqueous medium, (2) subjecting the said aqueous medium to continuous agitation during said slaking, (3) passing a suspension of calcium hydroxide obtained after slaking through a sieve having an aperture size of 40-70 microns, (4) subjecting the suspension to high energy high shear agitation with an impeller having a peripheral speed of 40-70 m/sec., so as to obtain finely dispersed calcium hydroxide, (5) terminating the said high energy high shear agitation on achieving finely dispersed slaked lime, (6) carbonating this suspension the finely dispersed slaked lime by passing therethrough sufficient gas comprising carbon dioxide to cause the neutralize the pH of the suspension to fall to neutral, during said carbonation step, (7) subjecting the said suspension to continuous agitation to maintain the suspension with an impeller speed of 200-700 cm/sec to maintain the suspension,- and separating (8) separating the precipitated calcium carbonate formed in the process. The disadvantage associated with this method is requirement of for generating high energy high shear agitation during slaking and carbonation.

The use of additives to control the morphology and particle size is recommended also reported. As per the US patent No. 5,558,850 (1996) to Bleakley Ian S. *et al.* (US patent 5,558,850) disclose, a process wherein 0.1 to 2.0% by weight of a reagent having one or more active hydrogen atoms e.g., polyhydric alcohol or phenol is added to the aqueous medium in which the quick lime is slaked. As per the US patent No. 5,332,564 (1994) to Chapnerkar Vasant D. *et al.* (US patent 5,332,564) disclose a process wherein, quicklime is slaked in an aqueous solution containing about 0.1 to 2.0% by weight of a sugar for the production of rhombic shaped precipitated calcium carbonate, and as per the US patent No. 5,232,678 (1993) to Bleakley Ian S. *et al.* (US Patent 5,232,678),

disclose a process wherein 0.01 to 1.5-% by weight of triethanolamine, mannitol, morpholine and solid boroheptonate are employed ~~in~~for the preparation of clusters of calcium carbonate which give good light scattering properties when used as a paper filler or paper coating pigment. The US patent No. 4,714,603 (1987) to

5 Vanderheiden, Denis B. (US patent 4,714,603) mention discloses the use of polyphosphates in an amount of 0.1 to 1.0 % by weight for getting generating precipitated calcite of substantially spherical morphology suitable for use in dull finish coated paper. The disadvantage associated with all these processes is the requirement of special reagents which adds to the production cost.

10 US Patent No. 5,833,747 (1998) to Bleakley Ian S. et al. (US Patent 5,833,747), discloses a method for preparing precipitated calcium carbonate for use as a pigment in paper coating compositions. The method comprising comprises the steps of (1) carbonating an aqueous medium containing lime, (2) at least partially dewatering the precipitated calcium carbonate—containing suspension using a pressure filter device operating at a pressure of 5 to 10 MPa, and (3) subjecting the precipitated calcium carbonate—containing suspension to comminution by high shear attrition grinding with an attrition grinding medium such as silica sand having a median particle diameter in the range 0.1 to 4.0 mm. The product predominantly comprises aragonitic or scalenohedral crystals. The disadvantages 15 of this method are that it required include the requirement of a device for high shear attrition grinding with a special grinding medium which is not separated during the process and a pressure filter device for dewatering the precipitated calcium carbonate containing suspension. Also, the grinding medium is not separated during the process.

20 25 The process as described in US Patent No. 5,695,733 (1997) to Kroc Vicki J. et al., disclose a process (US patent 5,695,733) that comprises the steps of (1) forming a reaction mixture containing seed material of a scalenohedral particles of aragonite type calcium carbonate and (2) adding lime slurry into the reaction mixture while simultaneously introducing carbon dioxide. The flow rates of the lime slurry and carbon dioxide are adjusted to control the solution conductivity of the reaction mixture to between from 2 to 4 milli Siemens to form the clusters of calcite particles. The drawbacks of this process is that it requires simultaneous

addition of lime slurry and carbon dioxide for maintaining the solution conductivity. Moreover, simultaneous control of flow rates of both liquid phase and gaseous phase reactants simultaneously is difficult to administer.

You Kyu Jae discloses A process for producing calcium carbonate particles having an average size of 0.1 to 1.0 micron is described in (US patent No.5,811,070 (1998) to You Kyu Jae. The process comprising comprises the following steps of, (1) introducing Carbon dioxide is introduced into a milk of lime containing a first reagent, such as consisting of sodium glutamate, sugar, and or a mixture thereof, in the range of 0.1 to 2.0 parts per 100 parts of calcium hydroxide initially present in the milk of lime, to prepare an aqueous suspension containing calcium carbonate particles of 0.4 micron in average size. The concentration of the first reagent is from 0.1 to 2.0 parts per 100 parts of calcium hydroxide initially present in the milk of lime, (2) adding a milk of lime is added into the above aqueous suspension, (3), and continuously reacting a carbonated solution is added to the aqueous suspension, which containing contains a second reagent including comprising at least one of sodium polyacrylate and a bicarbonate in the range of 0.1 to 5.0 parts per 100 parts of calcium hydroxide present initially, with the aqueous suspension. Calcium carbonate particles according to produced by the process are suitable as a filler for adhesives, paints, inks, papers and plastics, especially Polyethylene transparent polyethylene films. The drawbacks associated with this process is the addition of two different types of reagents in two stages, both of which are of different types, which. This makes the process complicated and unattractive.

US Patent No.4,367,207(1983) to Vanderheiden Dennis B discloses a process (US Patent 4,367,207), concerned with a process for preparing finely divided precipitated calcite. In the Vanderheiden process, in which carbon dioxide is introduced into an aqueous calcium hydroxide slurry containing anionic organo-polyphosphate polyelectrolyte with the carbonation started at a temperature above from about 7°C and below to about 18°C. The drawbaekOne disadvantage of this process is that it requires the requirement of an anionic polyelectrolyte which adds to the production cost. Another disadvantage is maintaining the required maintenance of a reaction temperature below ambient temperature. This

requirement necessitates, which require a chilling plant and it is which is energy consuming.

Summary of the Invention

5 The present invention provides a process for the production of precipitated calcium carbonate from calcium carbonate-rich by-product generated in a chemical processing industry which process obviates the drawbacks as detailed above. Precipitated calcium carbonate is produced by calcination of calcium carbonate-rich by-product into quick lime in a kiln. This is followed by slaking or hydration
10 to obtain hydrated lime. The hydrated lime is subsequently subjected to carbonation.

The present invention also develops an alternative source to limestone as a source for producing precipitated calcium carbonate useful for commercial applications.

15 The present invention also develops a process for purification, calcination, slaking and carbonation of a calcium carbonate-rich by-product having particle size in the range offrom 20 to 150 microns to produce precipitated calcium carbonate of particle size less than 20 microns.

20 The present invention also provides a continuous process for the purification of a calcium carbonate rich by-product from of a nitrophosphate fertilizer plant in order to thereby minimize minimizing solid waste produced by such a plant.

25 The present invention also provides pollution abatement measures in for a nitrophosphate fertilizer plant by utilization utilizing of the calcium carbonate-rich by-product generated in such a plant.

30 The present invention also produces high value finely divided precipitated calcium carbonate useful as filler in paints, plastics, rubber, poly vinyl chloride (PVC) and paper from calcium carbonate rich by-product generated in a nitrophosphate fertilizer plant, a high-value finely divided precipitated calcium carbonate which is useful as a filler in paints, and in plastics, rubber, poly vinyl chloride (PVC) and paper.

The present invention relates to a process for the production of precipitated calcium carbonate from calcium carbonate-rich by-product generated in industrial processes, specifically from a nitrophosphate plant or in the fertilizer industry. In one embodiment, the steps of the process comprise:

- 5 (1) calcination calcinating of a calcium carbonate rich by-product generated in a nitrophosphate fertilizer plant, the by-product having a moisture content up to 25% and a particle size in the range from 20 to 150 microns, in a rotary calciner at a continuous feed rate of from 5 to 20 kg/h at the calcination temperature of above 850°C and below about 950°C with the a residence time of from 60 to 90 minutes, so as to obtain a calcined material having from 75 to 88% available calcium oxide;
- 10 (2) removing water vapors, volatile matters, ammonia, NO_x and carbon dioxide during the calcination using a blower and a scrubber;
- 15 (3) slaking the calcined material (calcium oxide) in a slaker provided with an agitator rotating at 120 RPM to produce a hydrated lime slurry having a solids concentration in the range from 15 to 23% by weight;
- 20 (4) removing heavier and coarse particles from the hydrated lime slurry by wet sieving through a 60 to 100 mesh sieve to form a fine hydrated lime slurry;
- 25 (5) diluting the fine hydrated lime slurry to a desired solids concentration in the range from 10 to 20% by weight;
- 30 (6) taking transferring the diluted lime slurry to a carbonation tower and passing a carbon dioxide-air mixture containing 25 % by volume carbon dioxide at a superficial gas velocity of from 10 to 15 cm/sec. maintaining the at a maintained temperature in the range from 25 to 45°C until the pH of the diluted lime slurry fallen to is lowered to near neutral;
- (7) separating the precipitates formed in step (6) by known methods, e.g., filtration or centrifugation;
- (8) drying and pulverizing the solids separated precipitates to get produce a precipitated calcium carbonate; and
- (9) optionally, treating the product slurry before filtration the separation with a fatty acid or its salt or a salt of a fatty acid- such as, for example,

stearic acid or sodium stearate, at 95° C in the a concentration range of from 2 to 3.5% so as to obtain a coated precipitated calcium carbonate which is industrially useful in a number of applications such as rubber, plastics, paints and PVC.

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The invention is further directed to a precipitated calcium carbonate product having a mean particle size of from about 4 to 6 microns and wherein 100% of the particles are less than 20 microns and having a weight percent of calcium carbonate greater than 97.

10 The process described herein significantly purifies a calcium carbonate-rich by-product generated in a nitrophosphate fertilizer plant to get produce precipitated calcium carbonate.

Brief Description of the Figures

15 Figure 1 accompanying this specification represents the machines/ equipments of the unit apparatus used for the calcination and production of precipitated calcium carbonate from a calcium carbonate-rich by-product generated in a nitrophosphate fertilizer plant.

20 The unit apparatus embodies comprises the following equipmentscomponents:

M-paddle mixer cum screw feeder,

S₁-Indirect rotary dryer,

S₂-Indirect rotary calciner,

25 S₃- Jacketed Cooler,

P₁- Jaw crusher,

T₁-Slurry tank;

T₂ -lime slaker,

T₃ - sedimentation tank,

30 C- Carbonation tower with sparger,

F- Filtration unit, pressure or rotary vacuum type and

D-Flash dryer.

The numbers 1 to 9 in Fig. 1 indicate the following:

1. Calcium carbonate rich by-product, *i.e.*, Feed material.
2. Water
3. Coating agent solution
- 5 4. Steam
- 5 5. Filtrate
6. Product
7. Flue gases
8. Exhaust gases
- 10 9. Grit / sediments.

Detailed Description of the Invention

Accordingly, ~~t~~The present invention provides a process for the production of precipitated calcium carbonate from calcium carbonate-rich by-product generated in industrial processes, specifically from nitrophosphate fertilizer plants.
~~The process of the invention which comprises feeding continuously the a wet calcium carbonate-rich by-product containing up to 25% moisture and having particle size in the range offrom 20 to 150 microns, continuously into a calciner. The calciner is maintained in theat a temperature range from 850° to 950°C. The calciner is operated, with an angle of inclination of 1.08 degrees and a shell rotation speed of from 0.5 to 2 RPM. The calcium carbonate-rich by-product is added to the calciner, at a feed rate of from 5 to 20 kg/h, and The calciner is further provided with the arrangements means to remove the liberated water vapors, carbon dioxide, ammonia and NO_x-containing exhaust gas into a scrubber wherein ammonia and NO_x gases are scrubbed with water. The calciner is further provided with a means to, taking move the calcined outlet material eomning out exiting the calciner after residence time of about 30 to 90 minutes, and exiting the calciner at the outlet rate of 2.5 to 6 kg/h, to a crusher (e.g. a Jaw crusher)., The crusher erushing crushes the calcined material using a suitable crusher (e.g. Jaw crusher) to form small lumps. The crushed calcined material is, eooling cooled it into the a temperature range offrom 40 to 50°C, slaking and slaked with water in a slaker which is provided with an agitator, to produce a hydrated lime slurry having~~

a concentration from 15 to 23% solids by weight. The hydrated lime slurry is processed, followed by removing heavier and coarser particles by sedimentation and / or wet sieving through a 60 to 100 mesh sieve. The processed hydrated lime slurry, diluting the slurry is then diluted to a desired concentration in the range from 5 10 to 20% solids by weight. The diluted slurry is then taken, taking it up in a carbonation reactor and passing a carbon dioxide-air mixture is passed through the slurry. The carbon dioxide-air mixture containing contains 25% by vol. carbon dioxide and is passed through the slurry at the a gas velocity in the range from 10 to 15 cm/sec., maintaining the at a maintained temperature in the range off from 10 to 15 45°C. The carbon dioxide-air mixture is passed through the slurry, until the pH of the slurry has fallen to near neutral. Material that precipitates during the carbonation is then, separating separated the precipitated material formed by known methods, e.g. by filtration or centrifugation drying. The separated material is dried and pulverizing pulverized the solid to get produce precipitated calcium carbonate, optionally Optionally, treating the slurry is treated before filtration with a fatty acid derivative such as sodium stearate at 95°C in the a concentration range from 2.0 to 3.5% by weight for 10 minutes at 95°C. and allowing. The treated slurry is then allowed to cool to below 50°C with continuous stirring to get produce coated calcium carbonate. The slurry of coated calcium carbonate thus produced is 20 then separated followed by separation of the solid using known methods. The separated material is then drying dried and pulverizing pulverized the solid to get produce coated calcium carbonate.

In an one embodiment of the present invention, calcium carbonate-rich by-product having a moisture content up to 25% may be used as a starting material. 25 This provides the special advantage of using calcium carbonate-rich by-product as it is generated in the nitrophosphate fertilizer plant or as it is stored for a longer period in silos. Therefore, the process provided by the present invention can be linked to the nitrophosphate fertilizer plant. The process of the invention may thereby be made continuous with the production of calcium carbonate-rich by-product by the nitrophosphate fertilizer plant. and made continuous.

In another embodiment of the present invention, a calcium carbonate carbonate-rich by-product of a nitrophosphate fertilizer plant can be utilized on a

continuous basis to obtain quick lime, slaked lime and/or subsequently precipitated calcium carbonate for commercial applications.

In yet another embodiment of the present invention, the whiteness brightness, lightness, fineness, oil absorption, alkalinity and purity are improved by
5 the present invention.

In yet another embodiment of the present invention, moisture up to 25%, volatile matters from 7 to 8% and ammonium compounds from 0.1 to 0.3% (as ammonical Nitrogen) present in a calcium earbonate-carbonate-rich by-product is are removed.

10 In yet another embodiment of the present invention, nitrate impurities in the range from 0.05 to 0.1% (as nitrate Nitrogen) is-are removed from the-a calcium earbonate-carbonate-rich by-product generated in a fertilizer plant.

15 In yet another embodiment of the present invention, the material outlet rate from the calciner may be adjusted to-in a range from 2 to 6 kg/h by varying the rotation speed of the canciner from 0.5 to 2 RPM at the angle of inclination of 1.08°.

20 In yet another embodiment of the present invention, the calcination of calcium earbonate-carbonate-rich by-product may be carried out at the-a temperature above 850°C and below 950°C to obtain a calcined material having from 75 to 88 % available CaO.

In yet another embodiment of the present invention, the cooled and crushed calcined material may be slaked in water to obtain a hydrated lime slurry having from 15 to 23% solids in suspension.

25 In yet another embodiment of the present invention, the hydrated lime slurry may be diluted to a desired concentration in the range from 10 to 20% by weight solids and then wet sieved through a 60 to 100 mesh sieve to remove grit particles.

30 In yet another embodiment of the present invention, the hydrated lime slurry is subjected to carbonation in a bubble column reactor equipped with a suitable sparger and using a carbon dioxide-air mixture having a carbon dioxide concentration of 25% by volume.

In yet another embodiment of the present invention, the carbonation may be accomplished by passing the mixed gas at the-a rate of from 10 to 15 cm/sec and maintaining the temperature in the range from 25 to 45°C to achieve the-a neutral pH within from about 60 to 140 minutes.

5 In yet another embodiment of the present invention, the precipitated calcium carbonate formed in the process may be coated with a fatty acid derivatives to obtain from 2 to 3% total fatty matter-acid derivative in the final product.

10 In yet another embodiment of the present invention, the purity of the calcium carbonate-carbonate-rich by-product is improved from 88-90% to 97-99% by producing the precipitated calcium carbonate.

15 Calcium carbonate-carbonate-rich by-product generated in industrial processes, specifically in a nitrophosphate fertilizer plant, is-may containing comprise from 88- to 90% CaCO₃, from 0.2 to 0.4% total nitrogen-0.2-0.4%, from 0.1 to 0.3% ammonical nitrogen-0.1-0.3 %, from 0.05 to 0.1% nitrate nitrogen 0.05-0.1%, from 1.0 to 1.5% phosphate as P₂O₅-1.0-1.5%, from 0.2 to 0.25% Fluoride-fluoride as F-0.2-0.25%, from 3.5 to 4.0% mixed oxides (aluminum, iron, phosphates and matter insoluble in HCl)-3.5-4.0%, from 7 to 8% volatile matter-7-8%, from 43 to 45% loss on ignition-43-45%, from 0.1 to 0.15% soluble alkali as Na₂O-0.1-0.15 and from 1.0 to 1.2% acid insoluble matter-1.0-1.2%. MoreoverOther properties of the by-product include the following,- the The pH of a 5% aqueous suspension is from 9- to 9.3, the tapped bulk density is from 1.30- to 1.35 g/ml, the brightness/whiteness is 85% and the particle size in the range is from 20 to 150 microns.

25 Attempts were made to purify the-a calcium carbonate-carbonate-rich by-product by the methods known in the prior art. Physical methods such as sieving, floatation, sedimentation and hydrocyclone separation were not successful. which The failure of prior art processes led to believesuggested that the impurities present in the calcium carbonate-carbonate-rich by-product are an integral part of the material and not physically separate as in case of the conventional raw material,- limestone. Conventional chemical treatment methods were also found not suitable,

as because the chemicals used were unable to leach the entrapped impurities from the calcium carbonate-rich by-product.

Therefore, it was decided to calcine the calcium carbonate-carbonate-rich by-product to obtain calcium oxide. When the calcium carbonate-rich by-product was heated at to 900°C for one hour, the percent weight loss was 47.1% by weight and the X-Ray diffraction analysis indicated the presence of only calcium oxide. Therefore, it was inferred that some of the impurities can be removed by calcination of a calcium carbonate-carbonate-rich by-product, which by-product can be subjected to hydration and carbonation to obtain precipitated calcium carbonate.

Therefore, in the present invention, a process is developed for the production of precipitated calcium carbonate from a calcium carbonate-carbonate-rich by-product generated in industrial processes, specifically from a nitrophosphate fertilizer plant is developed, which The process comprises feeding continuously the a wet calcium carbonate-carbonate-rich by-product containing up to 25% moisture and having a particle size in the range of from 20 to 150 microns into a calciner maintained at above 850°C and below 950°C, at a feed rate of from 5 to 20 kg/h. The residence time of the material in the calciner is varied from 30 to 90 minutes by varying the rotation speed of the calciner from 0.5 to 2 RPM at an angle of inclination of 1.08 degrees. Calciner The calciner is provided with the arrangements means to remove the liberated water vapors, carbon dioxide, ammonia and NO_x-NO_x-containing exhaust gas into a scrubber, wherein ammonia and NO_x gases are scrubbed with water. The outlet rate of the calcined material varied from 2 to 10 kg/h. The calcined mass coming out is in a big lump form. Therefore, if the calcined material is passed through a jaw crusher fitted in-between the calciner and a material cooler. The crusher serves to crush the calcined material for crushing it into small lumps. The outlet material from the crusher is passed through water circulated jacketed a screw conveyer type material cooler equipped with a circulating water jacket. The material cooler serves for cooling the small lumps to near ambient temperature (from 40 to 50°C). Slaking of the calcined material obtained from the material cooler is carried out with water for one hour in a slaker provided with an agitator rotating which rotates at 120 RPM to

produce a hydrated lime slurry having a concentration from 15 to 23% by weight. Impurities present in the form of heavier and coarser particles were removed by sedimentation and / or by wet sieving through a 60 to 100 mesh sieve. The resulting slurry, from which the heavier/coarser particles have been removed, is
 5 diluted to a desired concentration in the range from 10 to 20% by weight and then transferred into a carbonation reactor provided with the a sparger. Carbonation is accomplished by passing a carbon dioxide-air mixture containing 25% by volume carbon dioxide, at a gas velocity of from 10 to 15 cm/sec. and maintaining the temperature in the range from 25 to 45°C until the pH of the slurry falls to near
 10 neutral, and separating. Then the precipitated material formed is separated by known methods. Drying and pulverizing. The separated solids are then dried and pulverized to get obtain precipitated calcium carbonate.

Optionally, the product slurry is was treated with fatty acid derivative an emulsion or solution of a fatty acid derivative at 95°C to get obtain from 2 to 3.5%
 15 by weight total fatty matter acid derivative in the a coated calcium carbonate. The solids were recovered by known methods of filtration, drying and pulverizing to get obtain the coated calcium carbonate. The properties of Precipitated-precipitated calcium carbonate and coated Precipitated-precipitated calcium carbonate are given in Table 1, which These properties indicated that it the products is would be
 20 useful as filler in paints, plastics, rubber and PVC.

Table 1. Properties of Precipitated and Coated calcium carbonate.

Sr.No.	Properties	Precipitated	Coated
1	CaCO ₃ , (by Acidimetry), % weight, Min.	97	95
2	pH (5% slurry)	10-11	9.5-10.5
3	Total Nitrogen, % weight	Not traceable	Not traceable
4	Ammonical Nitrogen, % weight	Not traceable	Not traceable
5	Nitrate Nitrogen, % weight	Not traceable	Not traceable
6	Phosphate as P ₂ O ₅ , % weight	less than 1.0	less than 0.1
7	Mixed oxides, % weight	2.0-3.0	1.50-2.5
8	Fluoride as F ⁻ , % weight	0.04	0.02
9	Bulk density, tapped, g/mlmL	0.40-0.45	0.35-.40
10	Oil absorption, % weight	60-100	80-100
11	Whiteness / Brightness , %	96-98	97-99
12	Particle size, microns	less than 20	less than 20
13	Total Fatty Matteracid derivative, % weight	NIL	2.5-3.0

The percent available CaO in the calcined material was determined by the Sugar method. This method involved adding a known weight of the sample in a 15% sugar solution and allowing it the sample to dissolve for 3 hours. The resulting solution was then filtering the solution filtered to remove undissolved undissolved solids, and titrating aAn aliquot of the filtrate was then titrated with a standardized acid.

5 The product properties were determined as per the by methods that are described as follows.

10 The calcium carbonate content of the product iswas determined by adding an accurately weighed about 2 g of sample (approximately 2 g) in to a known volume of standard acid, allowing itThe sample was allowed to react completely, and titrating the unused acidThe unreacted acid was then titrated with a standardized alkali solution. Mixed oxides arewere determined by precipitating hydroxides from an acidic solution of the sample, followed by ignition of The precipitated hydroxides were then ignited and weighing weighed as the corresponding oxides. Phosphates arewere determined by a spectrophotometric method. Ammonical nitrogen, nitrate nitrogen and fluoride arewere determined by ion-selective electrodes. Tapped bulk density iswas measured by noting the volume of material after tapping 50 times of a known weight of the sample 50 times in a calibrated cylinder. Whiteness/brightness iswas measured by measuring the reflectance of the sample and comparing with the reflectance of magnesium carbonate (considered as 100% whiteness) and with that the reflectance of carbon black (considered as zero percent whiteness). The particle size iswas determined using Mastersizer-2000 (Malvern, UK), particle size analyzer as on the dry powder.

15 20 25 The total content of fatty matter acid derivative in the coated precipitated calcium carbonate iswas determined by extracting the coated material with absolute alcohol, separating the same and evaporating the solvent to get isolate the coating material as a residue.

30 The process for the production of precipitated calcium carbonate according to the present invention, uses a new kind of raw material which is a by-product generated in a nitrophosphate fertilizer plant. This raw material has, hitherto not been used for the this purpose, This raw material having has a particle size in the

range from 20 to 150 microns and containing contains various types of impurities not usually found in the conventionally employed raw material, —limestone. Considering the scientific knowledge about the type of impurities present in a calcium carbonate-rich by-product, it was possible to find out the data for such a material with regard to its decomposition or sublimation temperature, which This information is described below:

Moisture can be removed from the calcium carbonate-rich by-product by drying the material at 105-110° C for a sufficient period. Ammonium carbonate decomposes at 58° C, ammonium bicarbonate decomposes between 10 36- and 60° C and sublimes, ammonium nitrate decomposes at 210° C, ammonium chloride and ammonium fluoride sublimate at 335° C.

Calcium nitrate (anhydrous) melts at 561° C, calcium nitrate trihydrate melts at 51.5° C, calcium nitrate tetrahydrate decomposes at 15 132° C. Calcium mono- and di-phosphates decompose below 205° C, whereas calcium tri- and pyro-phosphate melt above 1200° C. Calcium hydroxide is converted to CaO and water between 580 and 600° C.

Aragonite—Aragonite-type calcium carbonate decomposes at 825° C, whereas calcite type calcium carbonate decomposes at about 895° C, into Both 20 aragonite and calcite decompose to form Calcium oxide and Carbon carbon dioxide. If calcium carbonate is heated above 1000° C, it forms dead burnt / inactive Calcium oxide and liberates Carbon dioxide. It was-is also known that active calcium oxide could-can be slaked to obtain hydrated lime which can be further carbonated to produce precipitated calcium carbonate.

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- The advantages involved in the claimed process include:
- Using a raw material without pre-drying or crushing.
 - Keeping the material in free-flowing conditions throughout the processing.
 - Removal of the chemical impurities during calcination.
 - Improving brightness without using any special reagent.

- Calcination of very fine (from 20 to 150 micron) and highly dense (bulk density 1.3 to 1.35 g/mL) particles with minimum dusting.
- Avoiding the requirement of special types of material ~~of~~for construction of the calciner by limiting the calcination temperature to below 1000°C temperature.
- Limiting the effect of calcination to removal of impurities without forming the dead burnt lime.
- ArrangementsMeans for scrubbing the liberated gases to avoid pollution of the surrounding environment.

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The following examples are given by way of illustrations and therefore should not be construed to limit the scope of the present invention.

EXAMPLE -1

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The—A calcium carbonate—carbonate-rich by-product generated in a nitrophosphate fertilizer plant, having 12% moisture, was dried to a moisture content of <1% in a rotary dryer, ~~and~~The dried calcium carbonate-rich by-product was fed at the rate of 10 kg/h into a rotary calciner. ~~The rotary calciner was~~ preheated to 875°C with an angle of inclination 1.08 and a shell rotation speed of 20 1.5 RPM. The hold-up time was 35 minutes and the outlet rate was 5.2 kg/h. The product material was found to have 37% by wt. available CaO indicating incomplete calcination.

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EXAMPLE -2

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The—A calcium carbonate—carbonate-rich by-product generated in a nitrophosphate fertilizer plant, having 12% moisture was dried to a moisture content of <1% in a rotary dryer. ~~The dried calcium carbonate-rich by-product was~~ and fed at the rate of 10 kg/h into a rotary calciner. ~~The calciner was~~ preheated to 875°C with an angle of inclination of 1.08 degrees and a shell rotation speed of 30 0.75 RPM. The hold-hold-up time of the material in the calciner was 60 minutes and the outlet rate was 5.10 kg/h. The calcined mass obtained thereby was found to have 75% available CaO. The calcined material was crushed—after crushing—to into

small lumps. The crushed calcined material is~~was~~ subjected to slaking for a period of one hour in a slaker so as to obtain a suspension having about 20% by weight solids. If~~The suspension~~ was wet sieved through a 100 mesh sieve to remove grit particles. The sieved material was~~and~~ diluted to obtain a slurry containing about 5 10% solids~~slurry~~. This slurry is~~was~~ subjected to carbonation by passing through the slurry a carbon dioxide-air mixture containing 25% by volume carbon dioxide at a superficial gas velocity of about 12.5 cm/sec. and~~maintaining~~ the The temperature of the carbonation was maintained in the range from 25 to 45°C until the pH of the suspension falls~~dropped~~ to near neutral. The carbonation was 10 accomplished~~completed~~ within about 90 minutes. Thereafter, the product is~~was~~ recovered by known methods. The product was having~~contained~~_ about 80% calcium carbonate, with a bulk density of about 0.50 g/ml~~mL~~ and an average particle size of above 10 microns.

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EXAMPLE-3

The—A calcium carbonate-carbonate-rich by-product generated in a nitrophosphate fertilizer plant, having 12% moisture was dried to a moisture content of <1% in a rotary dryer. The dried calcium carbonate-rich by-product was~~and~~ fed at the rate of 5 kg/h into a rotary calciner. The calciner was preheated to 20 950°C with an angle of inclination of 1.08 degrees and a shell rotation speed of 0.50 RPM. The held~~hold-up~~ time was 90 minutes and the outlet rate was 1.7 kg/h. The calcined mass obtained was found to have 88% available CaO. The calcined material~~after crushing to~~was crushed into small lumps. The crushed calcined material is~~was~~ subjected to slaking for a period of one hour in a slaker so as to obtain a suspension having about 20% by weight solids. If~~The suspension~~ was wet sieved through a 100 mesh sieve to remove grit particles and diluted to obtain a slurry containing about 10% solids~~slurry~~. This slurry is~~was~~ subjected to carbonation by passing a carbon dioxide-air mixture containing 25% by volume carbon dioxide through the slurry at a superficial gas velocity of about 12.5 cm/sec. 25 and~~maintaining~~ the The temperature of the carbonation was maintained in the range from 25 to 45°C, until the pH of the suspension fallen~~dropped~~ to neutral. The carbonation was accomplished~~completed~~ within 90 minutes. Thereafter, the 30

product ~~is~~was recovered by known methods. The product ~~was having contained~~ about 98% calcium carbonate, with a bulk density of about 0.42 g/ml-mL and an average particle size in of about 6 microns with 100% of the particles having a particle size less than than 20 microns.

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EXAMPLE-4

The ~~A~~ calcium carbonate-carbonate-rich by-product generated in a nitrophosphate fertilizer plant, having 12% moisture was dried to a moisture content of <1% in a rotary dryer. The dried calcium carbonate-rich by-product was fed at the rate of 10 kg/h into a rotary calciner. The calciner was preheated to 950°C with an angle of inclination of 1.08 degrees and a shell rotation speed of 0.50 RPM. The hold-up time was 90 minutes and the outlet rate was 4.08 kg/h. The calcined mass obtained was found to have 88% available CaO. The calcined material was crushed into small lumps. ~~is~~The crushed calcined material was subjected to slaking after crushing to small lumps for a period of one hour in a slaker so as to obtain a suspension having about 20% by weight solids. ~~It~~The suspension was wet sieved through a 100 mesh sieve to remove grit particles and then diluted to obtain a slurry containing about 10% solids slurry. This slurry is was subjected to carbonation by passing a carbon dioxide-air mixture containing 25% by volume carbon dioxide through the slurry at a superficial gas velocity of about 12.5 cm/sec. and maintaining the The temperature of the carbonation was maintained in the range from 25 to 45°C, until the pH of the suspension fallen dropped to neutral. The carbonation was accomplished completed within 90 minutes. Thereafter, the product is~~was~~ recovered by known methods. The product ~~was having contained~~ about 98% calcium carbonate, with a bulk density of about 0.42 g/ml-mL and an average particle size of about 6 microns with 100% of the particles having a particle size of less than 20 microns.

EXAMPLE-5

30 The ~~A~~ calcium carbonate-carbonate-rich by-product generated in nitrophosphate fertilizer plant, having 23% moisture was dried to a moisture content of <1% in a rotary dryer. The dried calcium carbonate-rich by-product

was and fed at the rate of 10 kg/h into a rotary calciner. The calciner was preheated to 950° C with an angle of inclination of 1.08 degrees and a shell rotation speed of 0.50 RPM. The held hold-up time was 90 minutes and the outlet rate was 4.18 kg/hr. The calcined mass obtained was found to have 88% available CaO. The 5 calcined material was crushed into small lumps. The crushed calcined is material was subjected to slaking after crushing to small lumps for a period of one hour in a slaker so as to obtain a suspension having about 20% by weight solids. It The suspension was wet sieved through a 100 mesh sieve to remove grit particles and then diluted to obtain a slurry containing about 10% solids slurry. This slurry is 10 was subjected to carbonation by passing a carbon dioxide-air mixture containing 25% by volume carbon dioxide through the slurry at a superficial gas velocity of about 12.5 cm/sec. and maintaining the The temperature of the carbonation was maintained in the range from 25 to 45° C, until the pH of the suspension fallen 15 dropped to neutral. The carbonation was accomplished completed within 90 minutes. Thereafter, the product is was recovered by known methods. The product was having contained about 98% calcium carbonate, with a bulk density of about 0.42 g/ml mL and an average particle size of about 6 microns with 100% particles maximum particle size of less than 20 microns.

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EXAMPLE-6

Pre-dried calcium carbonate-carbonate-rich by-product having a particle size in the range from 20 to 150 microns and a moisture content of <1%, was introduced into a Fluidized Bed Calciner at the rate of 10 kg/h. A fluidized bed was formed by passing hot air through the calciner tower at a lower gas velocity of 25 about 100 cm/sec. The temperature of calcination was maintained at around 950° C by injecting fuel at a controlled rate. Within 60 minutes the desired calcined mass having 88 % available CaO is was obtained at the outlet rate of 5.7 kg/h. The calcined material is was subjected to slaking for a period of one hour in a slaker so as to obtain a suspension having about 20% by weight solids. It The suspension 30 was wet sieved through a 100 mesh sieve to remove grit particles and diluted to obtain a slurry containing about 10% solids slurry. This slurry is subjected to carbonation by passing through the slurry a carbon dioxide-air mixture containing

25% by volume carbon dioxide at a superficial gas velocity of about 12.5 cm/sec. and maintaining the The temperature was maintained in the range from 25 to 45°C, until the pH of the suspension fallen dropped to neutral. The carbonation was accomplished completed within 90 minutes. Thereafter, the product is was recovered by known methods. The product was having contained about 98% calcium carbonate, with a bulk density of about 0.42 g/ml and an average particle size of about 6 microns with 100% particles a maximum particle size of less than 20 microns.

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EXAMPLE-7

The A by-product calcium carbonate with 1% moisture was calcined at 950°C in a rotary calciner for 90 minutes to obtain 88% active calcium oxide. The calcined material which was crushed to small lumps using a jaw crusher and cooled to below 50°C using a water circulated jacketed screw conveyor conveyer-type material cooler. 7 Seven kg. of this material was added to 35 liters of water kept in a 50 liters capacity slaker. The slaker was, provided with a paddle type agitator which was operated at a maximum speed of 120 RPM to obtain a 20% weight by volume slurry. The reaction was found to be exothermic with a maximum rise in temperature of about 40°C. The slurry was continuously agitated for one hour so as to achieve complete conversion of calcium oxide to calcium hydroxide. This slurry is was subjected to carbonation by passing a carbon dioxide-air mixture containing 25% by volume carbon dioxide through the slurry at a superficial gas velocity of about 12.5 cm/sec. and maintaining t The temperature of the carbonation was maintained in the range from 25 to 45°C, until the pH of the suspension fallen dropped to neutral. The carbonation was accomplished completed within 90 minutes. Thereafter, the product is was recovered by known methods. The product was having contained about 98% calcium carbonate, with a bulk density of about 0.42 g/ml-mL and an average particle size of about 6 microns with 100% particles a maximum particle size of less than 20 microns.

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EXAMPLE-8

The-A by-product calcium carbonate with 1% moisture was calcined at 950°C in a rotary calciner for 90 minutes to obtain 88% active calcium oxide. The calcined material which was crushed to small lumps using a jaw crusher and cooled to below 50°C using a water circulated jacketed screw conveyor-type material cooler. It-The crushed calcined material was further pulverized to -150 mesh powder. The pulverized material (7 kg) of this material was added to 35 liters of water kept in a 50 liters capacity slaker, provided with a paddle type agitator. The slaker was operated at a maximum speed of 120 RPM to obtain a slurry containing 20% weight by volume slurry. The slurry was continuously agitated for one hour so as to achieve complete conversion of calcium oxide to calcium hydroxide. The reaction was found to be exothermic with maximum rise in temperature of about 40°C. This slurry is-was subjected to carbonation by passing a carbon dioxide-air mixture containing 25% by volume carbon dioxide through the slurry at a superficial gas velocity of about 12.5 cm/sec. and maintaining the temperature of the carbonation was maintained in the range from 25 to 45°C, until the pH of the suspension fallen dropped to neutral. The carbonation was accomplished completed within 90 minutes. Thereafter, the product is-was recovered by known methods. The product was having contained about 98% calcium carbonate, with a bulk density of about 0.42 g/mL and, an average particle size of about 6 microns with 100% particles and a maximum particle size of less than 20 microns. This indicated that there was no advantage in using a powdered calcined mass as compared to using a calcined mass crushed into small lumps in the slaking process.

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EXAMPLE-9

Seventy liters of the-a slaked lime slurry having about 20% solids, was prepared under similar conditions as similar to those described in Example-7. The slurry is-was transferred to a jacketed carbonation reactor consisting of a bubble column provided with a specially designed sparger specially designed to create sufficient turbulence for continuous mixing within the column. The reactor was further provided with arrangements for monitoring the reaction temperature and

the pH and cooling/heating arrangements using the jacket provided. The carbonation was accomplished completed within 90 minutes by passing mixture of air and a -CO₂-air mixture containing 25 volume % CO₂- at a superficial gas velocity of about 12.5 cm/sec. The carbonation temperature was maintained and 5 maintaining the temperature in the range from 25 to 45° C, wherein the mixed gas contained 25% by volume CO₂ and remainder being air. Sufficient A sufficient amount of the gas mixture was passed through the slurry to bring the pH of the slurry to neutral. When no further increase in pH (more than 7) was observed, the carbonation was stopped. Thereafter, The precipitated calcium carbonate formed in 10 the carbonation was isolated by filtration, followed by washing, drying and pulverizing. The product has contained about 98% calcium carbonate purity, with a bulk density in the range from 0.40-0.45 g/ml and, an average particle size of about 6 microns and a maximum particle size of with 100% particles below less than 20 microns.

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EXAMPLE-10

The A slurry of precipitated calcium carbonate obtained as per described in Example-7, was heated to 95° C under with continuous agitation. To this slurry was added; an emulsion of a fatty acid derivative such as stearic acid with caustic 20 soda, was added in an amount of 3.5% by weight of the calcium carbonate to be produced. The slurry was allowed to cool to 40-45° C under with stirring to get obtain a coated product. The coated product, which is was isolated by filtration, followed by washing, drying and pulverizing. The properties of this coated precipitated calcium carbonate is are as given provided in Table 1.

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EXAMPLE-11

The A slurry of precipitated calcium carbonate obtained as per described in Example-7, was heated to 95° C under with continuous agitation. To this slurry was added; an emulsion of a fatty acid derivative such as stearic acid with caustic 30 soda, was added in an amount of 2.0% by weight of the calcium carbonate to be produced. The resulting slurry was allowed to cool to 40-45° C under with stirring to get obtain a coated product. The coated product which is was isolated by

filtration, followed by washing, drying and pulverizing. The total content of total fatty matter-acid derivative in the coated product was 1.6% by weight. Other properties of the product were similar to those given provided in Table 1.

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EXAMPLE-12

The A by-product calcium carbonate with 1% moisture was calcined at 950°C in a rotary calciner for 90 minutes to obtain 88% active calcium oxide. The calcined material which was crushed to small lumps using a jaw crusher and cooled to below 50°C using water circulated jacketed screw conveyor-conveyer-type material cooler. 7 Seven kg. of this the crushed material was added to 35 liters of water kept in a 50 liters capacity slaker,. The slaker was provided with a paddle type agitator. The slaker was operated at maximum speed of 120 RPM to obtain a 20% weight by volume slurry. The reaction was found to be exothermic with maximum rise in temperature of about 40°C. The slurry was continuously agitated for one hour so as to achieve complete conversion of calcium oxide to calcium hydroxide. This slurry is-was subjected to carbonation by passing a carbon dioxide-air mixture containing 25% by volume carbon dioxide through the slurry at a superficial gas velocity of about 15.0 cm/sec. and maintaining the The carbonation temperature- was maintained in the range from 25 to 45°C, until the pH of the suspension fallen dropped to neutral. The carbonation was accomplished completed within 75 minutes. Thereafter, the product is-was recovered by known methods. The product was having contained about 97% calcium carbonate, with a bulk density of about 0.45 g/mlmL, and average particle size of about 7 microns and a maximum particle size of with 100% particles less then 25 microns.

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EXAMPLE-13

The A by-product calcium carbonate with 1% moisture was calcined at 950°C in a rotary calciner for 90 minutes to obtain 88% active calcium oxide. The calcined material which was crushed to small lumps using a jaw crusher. The crushed material and-was cooled to below 50°C using a water circulated jacketed screw conveyor-conveyer-type material cooler. 7 Seven kg. of this the crushed material was added to 35 liters of water kept in a 50 liters capacity slaker,. The

slaker was provided with a paddle type agitator. The slaker was operated at a maximum speed of 120 RPM to obtain a 20% weight by volume slurry. The reaction was found to be exothermic with maximum rise in temperature of about 40°C. The slurry was continuously agitated for one hour so as to achieve 5 maximum conversion of calcium oxide to calcium hydroxide. This The resulting slurry is-was subjected to carbonation by passing a carbon dioxide-air mixture containing 25% by volume carbon dioxide through the slurry at a superficial gas velocity of about 10.0 cm/sec. and maintaining + The carbonation temperature was maintained in the range from 25 to 45°C, until the pH of the suspension fallen 10 dropped to neutral. The carbonation was accomplished completed within 140 minutes. Thereafter, the product is-was recovered by known methods. The product was havingcontained about 97% calcium carbonate, with a bulk density of about 0.45 g/mlmL, and average particle size of about 7 microns and a maximum particle size ofwith 100% particles less than 25 microns.

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EXAMPLE-14

The-A slaked lime slurry was obtained as described in Example-7. The slaked lime slurry was diluted to 10% wt/vol. of CaO and taken in a continuous stirred tank reactor equipped with an agitator, having height/diameter ratio of about 20 2. Keeping +The slurry was kept continuously agitated at 100 RPM, a-A mixture of CO₂ and air in the-a ratio of 30 : 70 to 70 and a space velocity of 2 cm/sec was introduced at ambient temperature. The reaction pH was found to come to neutral towards the completion of carbonation within 50 minutes. On stopping When the gas flow was stopped the pH was increasingcontinued to rise up to 8.0. The 25 carbonation was continued for further 10 minutes to get obtain a steady pH of 7.0 for further 10 minutes. The-A solid product was isolated, dried and pulverized to get obtain a precipitated calcium carbonate and a coated product as described in Example-10. The properties of the coated precipitated calcium carbonate thus obtained was-were the same as given the properties provided in Table 1.

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The advantages of the present invention are include

1. Calcium carbonate-rich by-product generated in industrial processes, specifically in nitrophosphate fertilizer plants, is converted to precipitated calcium carbonate by calcination, slaking and carbonation using conventionally used machinery.
- 5 2. A calcium carbonate-rich by-product is converted to value value-added material suitable for applications in rubber, paint, PVC and plastics manufacture.
- 10 3. An alternative source to limestone is provided for producing precipitated calcium carbonate useful for commercial applications is found out disclosed.
- 15 4. A calcium carbonate-rich by-product having a fine particle size from 20 to 150 microns is successfully treated without contributing to pollution, due to dusting, which is commonly associated with such fine powders due to dusting.
5. A continuous process has been developed is provided according to the present invention, which can be linked with to the plant generating the calcium carbonate-rich by-product. Thus, This linked process can minimized the solid waste pollution caused by such by-products and improved the environmental quality of a nitrophosphate fertilizer plant.
- 20 6. As Because the temperature of calcination of the a calcium carbonate-rich by-product is around 950°C, special types of material of construction materials is are not required to be used in the calciner.
7. Calcination according to the present invention provides the special advantage of removing all the major impurities simultaneously and converting the calcium carbonate-rich by-product to active calcium oxide. Therefore, not many few unit operations are required in claimed purification process.

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PROCESS FOR GENERATION OF PRECIPITATED CALCIUM
CARBONATE FROM CALCIUM CARBONATE
RICH INDUSTRIAL BY-PRODUCT

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ABSTRACT OF THE INVENTION

An environmentally friendly process for the production of precipitated calcium carbonate suitable for industrial application from a calcium carbonate-carbonate-rich by-product of a nitrophosphate fertilizer plant. by adopting The process 15 employs purification steps of calcination, carbonation and optional treatment with an emulsion of fatty acid or its fatty acid derivatives.